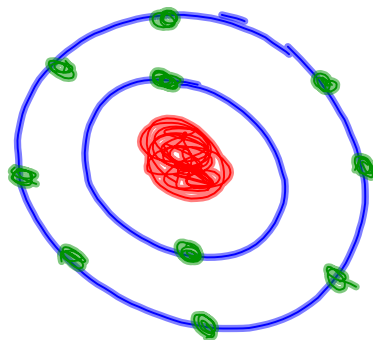


A Planetary Model of the Atom

Niels Bohr, suggested the following:

- electrons can move around the nucleus in nearly circular orbits
- each electron has a specific amount of energy
- the farther away from the nucleus the greater the amount of energy
- electrons cannot exist 'between' these orbits, but can move up and down from one orbit to another
- the order of filling these orbits is 2, 8, 8 for the first three orbits
- electrons are more stable at lower energy, closer to the nucleus

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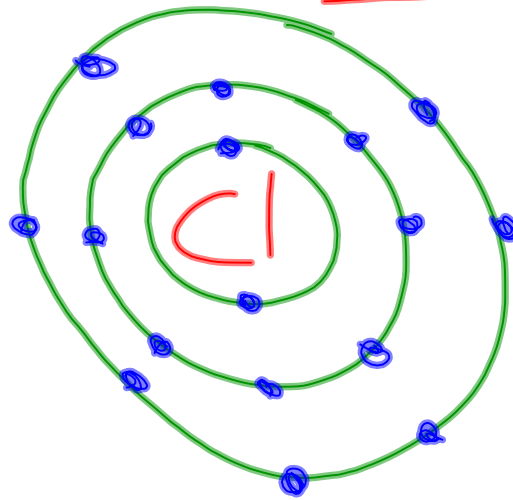


10 p
10 e⁻
Ne

Feb 21-9:11 AM

Atomic 17

Bohr Diagram



Feb 21-9:13 AM

Bohr Diagrams the element symbol is written in the center and the electrons are 'filled' into the orbits around this nucleus

Example:
Hydrogen has 1 proton, and 1 electron

● Electron

H

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Nitrogen has 7 protons and 7 electrons
(Hint: remember 2,8,8)

N



Mar 25-7:50 PM

Try these!
determine the # of protons and electrons
Dont forget the 2, 8, 8

● Electron

P

Cl

Mg

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