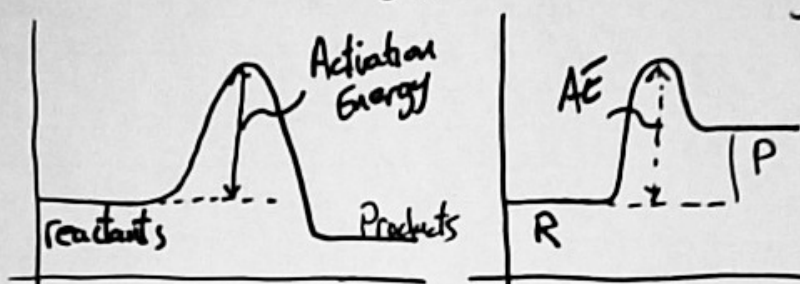


## Reaction Rates / Equilibrium - Chapt 12

### - Activation Energy & Collision Theory



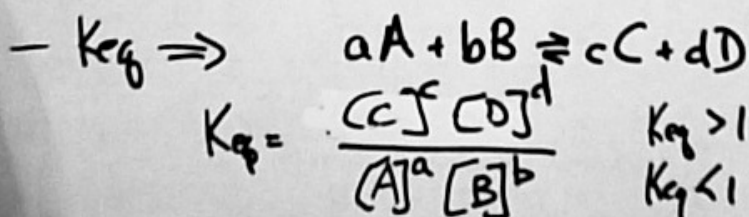
### - Factors Affecting Reaction Rates

- Temp, Concentration, Part. Size, Catalyst

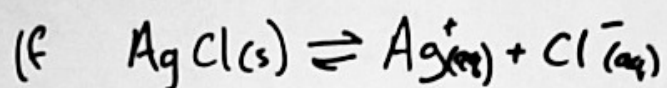
### - Reversible Rxn & Equilibrium

↳ Le Chateliers Principle

↳ Concentration, Temp, Pressure



### Solubility Equilibrium



$$K_{sp} = [\text{Ag}^+][\text{Cl}^-] \quad \because [\text{AgCl}] \Rightarrow \text{ignored.}$$

∴

If  $[x] > K_{sp}$  for X, ppt occurs

if  $[x] < K_{sp}$  for X, No ppt!